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Note: Not all of the reactions will occur.
For those that do not, write no reaction.

1. A piece of copper is dropped into a
container of water. No reaction
2. Liquid bromine is added to a container of

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sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)}$
 $2\text{NaBr(s)} + \text{I}_2\text{(s)}$ 3. An aluminum strip is
immersed in a solution of silver nitrate.

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complex the molecule, have an oxidation number (valence or charge) equal to zero. O₂, Fe, T₂, He < /1 1 ~ ~ L-td. 8. The following are diatomic or polyatomic elements in nature which must be committed to memory.

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hydrogen carbonate NaHCO 1. 2. 3. 4. 5.
6. 8. 9. 10. 11. vanadium(V) oxide H₂O

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dihydrogen monoxide (NH) CO
ammonium oxalate 4 224 polonium(VI)
thiocyanate Po(SCN)₆ tetraphosphorus
decaoxide P₄O₁₀ 12. 13. 15. 16. 17.
18. 19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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part (i) write a BALANCED equation and
in part (ii) answer the question about the
reaction.

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The Ultimate Chemical Equations
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number of the underlined element:

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KMnO₄ Since K is an alkali metal, its charge must be 1+ Oxygen is 2— but there are four of them, therefore, 4 times 2— equals 8— If 1+ and 8— are added Ultimate Chemical Equation Answers

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The Ultimate Chemical Equations
Handbook Determine the oxidation
number of the underlined element:
KMnO₄ Since K is an alkali metal, its
charge must be 1+ Oxygen is 2— but
there are four of them, therefore, 4

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times 2— equals 8— If 1+ and 8— are added The Ultimate Chemical Equations Handbook Answers Chapter 10

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