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CHAPTER 17, Thermochemistry

(continued) 2. What is calorimetry?

Calorimetry is the accurate and precise measurement of heat change for chemical and physical processes. 3. Use Figure 17.5 on page 511. Circle the letter next to each sentence that is true about calorimeters. G The calorimeter container is insulated to minimize loss of heat to or

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Thermochemistry183 SECTION 17.1 THE  
FLOW OF ENERGY—HEAT AND WORK  
(pages 505–510) This section explains  
the relationship between energy and  
heat, and ...

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Chapter 17 Thermochemistry 187 10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide.  
SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520–526) This section explains heat transfers that occur during melting, freezing, boiling, and condensing.

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## **SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505-510)**

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Answers 2. 0.100 g of H<sub>2</sub> and an excess  
of O<sub>2</sub> are compressed into a bomb  
calorimeter containing 1200 g of water.  
The temperature before the reaction is  
25.00 °C, and after the reaction it goes  
to 27.16 °C. The Page 2/15. Bookmark  
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Date \_\_\_\_\_ Class \_\_\_\_\_ CHAPTER 17,

Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the point of view of the system, or of the

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Unit 9: Thermochemistry Homework Packet (70 points)

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Name:      KEY      Period:      By the end of Unit 9, you should be able to:

Chapter 17 13. state the laws of thermodynamics 14. state the difference

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between temperature and heat 15.  
identify variables and their labels for  
heat quantities and solve

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of O<sub>2</sub> are compressed into a bomb  
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